Original Article

Growth and characterization of Sm\textsuperscript{3+} doped cerium oxalate single crystals

Minu Mary C\textsuperscript{a}, Vimal G\textsuperscript{a}, Kamal P. Mani\textsuperscript{a}, Gijo Jose\textsuperscript{b}, Biju P.R.\textsuperscript{a}, Cyriac Joseph\textsuperscript{a,∗}, Unnikrishnan N.V.\textsuperscript{a}, Ittyachen M.A.\textsuperscript{a}

\textsuperscript{a} School of Pure and Applied Physics, Mahatma Gandhi University, Kottayam, India
\textsuperscript{b} Department of Physics, S B College, Changanassery, India

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\section*{Abstract}

Single crystals of Sm\textsuperscript{3+} doped cerium oxalate decahydrate were synthesized using single diffusion gel technique and the conditions influencing the size, morphology, nucleation density and quality of the crystals were optimized. Highly transparent single crystals of average size 3 mm × 2 mm × 1 mm with well-defined hexagonal morphology were grown during a time period of two weeks. X-ray powder diffraction analysis revealed that the grown crystals crystallize in the monoclinic system with space group P2\textsubscript{1}/c as identical with the pure cerium oxalate. The various functional groups of the oxalate ligand and the water of crystallization were identified by Fourier transform infrared spectroscopy. The photoluminescence spectrum of the Sm\textsuperscript{3+} doped cerium oxide indicates that the Sm\textsuperscript{3+} ions are optically active in the cerium oxide matrix. The crystal has a strong and efficient orange red emission with a wavelength peak at 595 nm and hence can be effectively used for optical amplification. Microhardness measurements of the crystal revealed that they belong to the soft material category.

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\section*{1. Introduction}

High quality, defect free single crystals have always evoked great attention due to their wide applications in the field of optoelectronics, solid state lasers, remote sensing and medical diagnostics \cite{1-3}. Particular attention can be given to single crystals of trivalent lanthanide compounds owing to their smart optical and spectroscopic properties, which led to potential applications in optoelectronics, lasers, optical amplifiers and in telecommunications \cite{4-6}. The distinctive properties originate from the electronic transitions within the 4f shell of the rare earth ion that is well shielded by 5s and 5p electrons but highly sensitive to the characteristics of the host lattice. Among the rare earth compounds, rare earth oxalates are of special importance because of their interesting luminescent, magnetic and electrical properties \cite{7-10}. Recent investigations on the fluorescence of some rare earth oxalates suggest their potentiality for optical applications \cite{7,11}. Oxalates are also utilized as a precursor for many scientific researches.
of the technologically important ferroelectric, magnetic and superconducting materials such as BaTiO₃, Ni–Co–Zn ferrites and YBa₂Cu₃O₇−ₓ. Significant attention has been paid to rare earth ion doped crystals, since the dopants play a vital role in enhancing their physical properties. Of the various rare earth ions, Sm³⁺ ions have gained particular attention and are widely investigated in a variety of systems such as bulk crystals, nanocrystals and glasses, due to their high fluorescence quantum yield [5]. Hence, Sm³⁺ ion doped rare earth oxalates deserve special attention and importance.

Here we describe the growth of samarium doped cerium oxalate decahydrate single crystals. Gel method is the only viable method to grow these crystals since oxalates are sparingly soluble in water and decompose before melting, which impose constraints on the use of other conventional techniques. In addition, gel method also allows ease optimization of physical parameters influencing the growth of high quality single crystals. The present work mainly aims to investigate the growth parameters and the effect of Sm³⁺ doping on the optical properties of the cerium oxalate single crystal. The structural and mechanical properties of the crystal are also studied in detail.

2. Experimental

2.1. Crystal growth

Cerium nitrate hexahydrate (Ce(NO₃)₃·6H₂O, 99.9%, CDH), Samarium nitrate hexahydrate (Sm(NO₃)₃·6H₂O, 99.9%, CDH), oxalic acid dihydrate (H₂C₂O₄·2H₂O, 99%, Merck) and sodium meta silicate nonahydrate (Na₂SiO₃·9H₂O, CDH) were used as the starting materials. Sodium meta silicate solution of density 1.03 g/cc was mixed with 1 M oxalic acid to obtain hydrox silica gel of pH 6. The solution was transferred to a test tube of dimension 15 mm in diameter and 150 mm in length to fill half of its volume and then kept undisturbed for 24 h for proper gelation. 0.5 M aqueous solutions of cerium nitrate and samarium nitrate were mixed in stoichiometric proportions and poured gently on the top of the gel column. The reactants slowly diffuse through the gel and react with the oxalic acid already incorporated in the medium. The chemical reaction involved is

\[(2 - x)\text{Ce(NO₃)₃} · 6\text{H₂O} + x\text{Sm(NO₃)₃} · 6\text{H₂O} + 3\text{H₂C₂O₄·2H₂O} \rightarrow \text{Ce}_2\text{Sm}_x\text{(C₂O₄)₃·10H₂O} + 6\text{HNO₃} + 8\text{H₂O}\]

where \(x\) is taken as 0 and 0.1 for getting pure and 5% Sm doped cerium oxalate decahydrate, respectively. It is observed that whatever be the pH and density of the gel, it cracked as soon as the rare earth solutions were added over the set gel. The concentration of the feed solution was then varied but the observations were the same. To overcome these difficulties neutral gel and double diffusion systems were also undertaken. But the result was not as expected.

Finally, the single diffusion technique was again tried with the feed solution acidified with dilute nitric acid in various proportions. The observation revealed that the presence of nitric acid in the feed solution plays an important role in defining the nature of the precipitation. The density of the precipitation front lowers as the amount of HNO₃ in the feed solution increases. When concentrated HNO₃ was used to acidify the feed solution, there was no precipitation but spherulitic crystals are formed. As the volume of concentrated HNO₃ was further increased, tiny well faceted crystals were obtained. Hence, the acidification of the feed solution was found to be a necessary step in the growth of these groups of crystals.

Morphological changes are exhibited by the crystals due to the variation of the microenvironments like density of the gel, pH of gel, acidity of feed solution and degree of supersaturation. It is observed that crystals taken from different regions of the gel medium show changes in morphology. Fig. 1(a) and (b) shows growth system of Sm³⁺ doped cerium oxalate crystals in silica gel. It is observed that the transparent crystals formed at the bottom of the gel column are thin and have most developed hexagonal prism face while the crystals found at the top of the gel column are of thick rhombohedral prism morphology. Good quality crystals are found at the bottom of gel column since the availability of the ions is just sufficient for the growth of crystals. Distance of crystallization zone from the gel interface is also different with the acidity of the feed solution. With lower concentration of acid, the heavy nucleation resulted especially at the gel–solution interface as shown in Fig. 1(a). As the acid concentration is increased, nucleation density is decreased and crystallization occurred in a region below gel solution interface, resulting the formation of transparent hexagonal crystals as shown in Fig. 1(b). High concentration of the reactants at the interface of the feed solution and gel surface leads to spontaneous multinucleation resulting to the formation of clustered crystals as shown in Fig. 1(c) and (d). Fig. 1(e) shows the well-developed hexagonal crystal of average size 4 mm × 2 mm × 1 mm grown in the system. Clustered crystal of Sm³⁺ doped cerium oxalate is given in Fig. 1(f) and reveals that these type of crystals are formed by the intergrowth of two or more individual crystals in such a way that they grow together with only part of their similar faces visible.

The effects of the various growth parameters such as density of SMS (sodium meta silicate) solution, pH of the gel, concentration of the reactant and acidity of the feed solution on growth process are manipulated to optimize the conditions of growing good quality crystals. For getting maximum sized, good quality crystals, the optimized conditions are:

- Density of the SMS solution – 1.03 g/cc
- pH of the gel – 6
- Aging of the gel–24 h
- Concentration of inner reactant – 1 M
- Concentration of the feed solution – 0.5 M
- Acidity of the feed solution – 50% by volume, pH < 2

2.2. Characterization

The structure of the grown crystals was identified by X-ray powder diffraction analysis in the range 10–40° using PANalytical X’Pert Pro X-ray diffractometer with Cu-Kα radiation operating at 30 mA, 40 kV. FTIR absorption spectrum of the crystal was recorded using Shimadzu 8400S FTIR spectrometer in the range 400–4000 cm⁻¹. The absorption and emission spectra of the crystals were recorded using Shimadzu UV Probe
Fig. 1 – (a, b) Growth system of Sm³⁺ doped cerium oxalate crystals with different acid concentration, (c, d) growth system with clustered morphology, (e, f) morphology of a Sm³⁺ doped cerium oxalate single crystal and clustered crystal, respectively.

spectrophotometer in the wavelength range 250–700 nm and Shimadzu RFPC 5301 spectrofluorophotometer in the range 525–725 nm, respectively. Microhardness of the crystals was measured using a Vicker's microhardness tester.

3. Results and discussion

3.1. X-ray diffraction analysis

The XRD patterns of pure Ce₂(C₂O₄)₃·10H₂O and Sm³⁺ doped Ce₂(C₂O₄)₃·10H₂O single crystals were shown in Fig. 2. All diffraction peaks are matching well with the monoclinic structure of cerium oxalate decahydrate having lattice parameters a = 11.34 Å, b = 9.630 Å, c = 10.392 Å and β = 114.5° (ICPDS Card No 20-0268) and comparable with the lattice parameters of the Ln₂(C₂O₄)₃·10H₂O as reported by Sheng-Hua et al. [15]. Hence, it may be suggested that the grown crystals exhibit monoclinic structure with space group P2₁/c in the decadydrate form. According to Sheng-Hua et al. the structure is composed of two-dimensional networks of edge sharing 1:5:3 coordination polyhedral matching the (020) set of planes. Each lanthanum atom being surrounded by three chelating oxalate groups and three aqua ligands. The intervening space is filled by lattice water molecules disordered over seven major sites. The identical chemical constitution and crystallographic parameters of rare earth oxalates point to the fact that doped or mixed crystal of any two of them can form by the simple substitutional exchange of rare earth ions. Since the ionic radius and chemical nature of cerium and samarium are comparable with the above-mentioned rare earth elements, one can
The recorded crystals deformation $C_{IR}$ of cerium oxalate single crystal. FTIR spectrum of cerium doped Sm$^{3+}$ oxalate single crystal. The atomic percentage of the Ce and Sm is measured to be 13.18 and 0.83, respectively, which supports the doping concentration of Sm$^{3+}$.

### 3.4. Optical properties

#### 3.4.1. UV–visible absorption

The optical absorption spectrum of pure and samarium doped Ce$_2$(C$_2$O$_4$)$_3$·10H$_2$O crystals are given in Fig. 5(a) and (b), respectively. The broad absorption observed in both the spectra in the UV region is due to the transition from 4f to 5d of Ce$^{3+}$ ions. Sm$^{3+}$ ion will exhibit several overlapped excitation bands from 360 to 500 nm due to its several closely spaced energy levels located in this region [11]. Hence the spectrum of Sm$^{3+}$ doped cerium oxide exhibits sharp absorption peaks at nUV and visible regions. The transitions in the spectrum are originated from ground state 6$^2$H$_{5/2}$ to various multiples of 4$^f$ configurations of Sm$^{3+}$. The major absorption peaks of Sm$^{3+}$ observed in the spectrum at the wavelength 402, 417, 442, 463, and 479 nm are assigned to the transitions from ground state 6$^2$H$_{5/2}$ to the excited states $^2$F$_{7/2}$, $^2$P$_{3/2}$, $^4$G$_{9/2}$, $^4$I$_{13/2}$ and $^4$I$_{11/2}$ of Sm$^{3+}$, respectively [11]. A more enlarged view of the absorption peaks is given in the inset of Fig. 5(b).

#### 3.4.2. Photoluminescence studies

The emission spectrum of the pure and Sm$^{3+}$ doped cerium oxide single crystals when excited at 401 nm is given in Fig. 6(a) and (b). The broad blue-green emission band around 400–500 nm observed in both spectra is originated by the allowed electric dipole transitions ($^2$D$_{5/2}$, $^2$D$_{3/2}$)5d $\rightarrow$ 4f($^2$F$_{7/2}$, $^2$F$_{5/2}$) of Ce$^{3+}$ ions. Besides the emission of Ce$^{3+}$ ions, spectrum of Sm$^{3+}$ doped cerium oxide exhibits sharp emissions at 560, 595 and 643 nm due to the 4f–4f transitions of Sm$^{3+}$ ions. On exciting at 401 nm, the $^4$F$_{7/2}$ level of Sm$^{3+}$ become excited and after that the electrons nonradiatively relaxes to the $^4$G$_{5/2}$ level by multiphonon relaxation. Between the $^4$F$_{7/2}$ and $^4$G$_{5/2}$ levels there are number of closely spaced higher energy levels that ensure the fast nonradiative decay. The level $^4$G$_{5/2}$ has

### 3.2. FTIR analysis

FTIR absorption spectrum of the grown samarium doped cerium oxide single crystals was given in Fig. 3. The broad envelope extending from 2800 to 3600 cm$^{-1}$ is assigned to be due to the symmetric and asymmetric stretching modes of the water molecules. The strong band appearing in the IR spectrum around 1615 cm$^{-1}$ can be identified as due to the asymmetric stretching vibrations of C–O groups of the C$_2$O$_4^{2-}$ ions together with the bending mode of water. The strong peak around 1316 cm$^{-1}$ is also assigned to the asymmetric stretching of C–O groups. The strong bands around 495 and 796 cm$^{-1}$ are due to the combined effect of the in-plane deformation mode O–C–O and M–O bond and the weak one around 582 cm$^{-1}$ observed in the spectrum represent the wagging mode [11,16,17]. The FTIR spectroscopic analysis of single crystals of Ce$_{2-x}$Sm$_x$(C$_2$O$_4$)$_3$·10H$_2$O confirmed the presence of functional groups associated with the oxalate ligands and the metal–oxygen bond.

### 3.3. EDS analysis

The EDS pattern of Sm$^{3+}$ doped Ce$_2$(C$_2$O$_4$)$_3$·10H$_2$O crystal was recorded and is shown in Fig. 4. The EDS spectrum confirmed the presence of cerium and samarium in Sm$^{3+}$ doped cerium oxide single crystal.

![Fig. 3 – Fourier transform infrared spectrum of Sm$^{3+}$ doped cerium oxalate single crystal.](image)

![Fig. 4 – Energy dispersive spectrum of Sm$^{3+}$ doped cerium oxalate single crystal.](image)
sufficient energy gap higher than five phonons with respect to the next lower level. This encourages radiative transition to the ground state. Hence the emission spectrum of Sm\textsuperscript{3+} doped cerium oxalate exhibits three emission transitions from the excited level \(6\text{G}_{7/2}\) of Sm\textsuperscript{3+} ions to lower ground levels \(6\text{H}_{5/2}\), \(6\text{H}_{7/2}\) and \(6\text{H}_{9/2}\) corresponds to 560 nm, 595 nm and 643 nm, respectively \[5\]. The emission spectrum reveals that the emission from \(5\text{G}_{7/2}\) to \(6\text{H}_{7/2}\) at 595 nm has maximum intensity and resulting strong emission in the orange region.

### 3.5. Microhardness measurements

Perfectly plane single crystals of Sm\textsuperscript{3+} doped cerium oxalate were selected for the microhardness measurements. Indentation was carried out, using a Vickers microhardness tester (Leitz Wetzlar) for various loads ranging from 5 to 100 g. The duration of the indentation time was kept constant at 30 s and in all measurements, the distance between the two indentations was kept 10 times greater than the diagonal length of the indentation mark in order to avoid any mutual influence of the indentations. For each load, several indentations were done to calculate the microhardness. The hardness number \(H_v\) of the crystal was calculated using the equation:

\[
H_v = \frac{1.854P}{d^2} \text{ kg/mm}^2
\]  

The increase in hardness with load is primarily caused by the work hardening of the surface layers. Fig. 7 shows the variation of microhardness \(H_v\) against applied load in the case of Ce\textsubscript{2-x}Sm\textsubscript{x}(C\textsubscript{2}O\textsubscript{4})\textsubscript{3}·10H\textsubscript{2}O crystals. It was found that as load increases, microhardness increases initially and then due to the rearrangements of dislocations and mutual interactions between them, the value of microhardness becomes almost constant above a particular load.

Inset of Fig. 7 shows the variation of \(\log P\) vs \(\log d\). Meyer’s index number was calculated from Meyer’s law, which relates the load and indentation diagonal length.

\[
P = kd^n
\]


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